

## THE GAS STATE – LESSON PROGRESSION

CHAPTER	KEY TERMS	ACTIVITY	HOME WORK
9.1	Kinetic Molecular Theory States of Matter Solid, Liquid, gas	Gas Occupies Space, Bottle Balloon, Empty Pop can and Ice.....	Page 422 # 1 to 4
9.2	Boyles Law $P \propto 1/V$ $PV = \text{Constant}$ $P_1V_1 = P_2V_2$ Pressure Atmospheric Pressure Units of Pressure Pascal, Atmosphere, mm of Hg, torr	Lab using Pressure sensor (if the system is working) Graphical representation	Page 428 # 5 to 10
9.22	Charles Law $V \propto T$ Or $V/T$ is a constant $V_1/T_1 = V_2/T_2$ Absolute scale or Kelvin Scale $-273.14 = 0K$ Combined Gas Law $PV = nRT$ Ideal Gas Equation (9.4) $P_1V_1/T_1 = P_2V_2/T_2$		Page 423 # 16 to 19  Page 438 # 26 to 33
9.3	Compressed Gases		Read Pages 441 to 442 # 1 to 2
9.4	Ideal Gas equation $PV = nRT$		Pages 445 # 1 to 6
9.5	Air Quality Nitrogen Cycle Ground Level Ozone		Page 543 # 1 to 6

CHAPTER	KEY TERMS		HOME WORK
10	Dalton's Law of Partial Pressures Partial Pressure		Q # 1 Page 461 Q # 5,6,7 Pg 463
	Vapour Pressure Vapour Pressure of Water		Pg 465 Q # 8, 9
	Law of Combining Volumes or Gay Lussac's law Avogadro's Hypothesis or Theory		Pg 468 Q # 1,2,3,4
	Molar Volume		Pg 471 Q # 7, 9,11
	Montreal Protocol CFC's		Read Pages
	Ozone Layer		Read Pages 475 to 479 475 to 479 Q# 1,2,3